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Chemistry IF8766 .

MOLARITY BY DILUTION

Name Acids are usually
acquired from chemical
supply houses in
concentrated form.

These acids are diluted
to the desired

concentration by

adding water. Since

moles of acid before

dilution = moles of acid

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after dilution, and
moles of acid = $M \times V$
then,

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Molarity is expressed in
units of moles per liter
(mol/L). It's such a
common unit, it has its
own symbol, which is a
capital letter M. A

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solution that has the concentration 5 mol/L would be called a 5 M solution or said to have a concentration value of 5 molar.

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where the subscripts
"1" and "2" refer to the
solution before and

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after the dilution, respectively. Since the dilution process does not change the amount of solute in the solution, $n_1 = n_2$.

Thus, these two equations may be set equal to one another:

$$M_1V_1 = M_2V_2$$

This relation is commonly referred to as the dilution equation.

4.5: Molarity and Dilutions - Chemistry

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add to achieve a
specified volume and

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concentration. The calculator uses the formula $M_1 V_1 = M_2 V_2$ where "1" represents the concentrated conditions (i.e. stock solution Molarity and volume) and "2" represents the diluted ...

Solution Dilution Calculator | Sigma- Aldrich

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molarity and normality
are measures of
concentration. One is a
measure of the number
of moles per liter of
solution, while the
other is variable,

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Molarity is expressed in

units of moles per liter

(mol/L). It's such a

common unit, it has its

own symbol, which is a

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capital letter M. A solution that has the concentration 5 mol/L would be called a 5 M solution or said to have a concentration value of 5 molar.

Molarity Definition as Used in Chemistry

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**Dilution Problems,
Chemistry, Molarity**

& Concentration

Examples, Formula

& Equations

This chemistry video

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tutorial focuses on molarity and dilution problems. It shows you how to convert between molarity, grams, moles, and liters. It's very useful for students learning solution ...

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