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shows the lowest whole # ratio of elements in a.... a formula that shows the true ratio of elements in a compound.

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a positive or negative charge resulting  
from the loss or gain of electrons,  
respectively.

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Naming Ions & Counting Atoms: Chapter

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Section 9.4 – Naming and Writing Formulas for Acids and Bases. An acid is a compound that produces  $H^+$  ions when it dissolves in water. The formula for an acid normally starts with an H. When naming acids, you should first determine whether or not the acid

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contains oxygen.

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Names and Formulas 85 Names and Formulas for Bases (page 273) 6. A base is a compound that produces \_\_\_\_\_ when dissolved in water. 7. How are bases named?

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Formulas83 SECTION 9.3 NAMING AND WRITING FORMULAS FOR MOLECULAR COMPOUNDS (pages 268–270) This section explains the rules for naming and writing formulas for binary molecular compounds. Naming Binary Molecular Compounds (pages 268–269)

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CHAPTER 9 Assessment 42. a. 2- b. 1+

c. 1- d. 3+ 43. a. 2+ b. 2+ c. 3+ d. 1+

44. a. barium ion b. iodide ion c. silver

ion d. mercury(II) ion 45. cyanide, CN-

and hydroxide, OH- 46. a. hydroxide ion

b. lead(IV) ion c. sulfate ion d. oxide ion

47. zero; A compound is electrically

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neutral. 48. The symbols for the cation and

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are always in the same proportions any  
atom or group of atoms that has a  
negative charge ... place cation name  
first followed by the anion name: naming  
polyatomic ...

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Chapter 9 Chemical Names and  
Formulas 221 Name Date 4 Write  
formulas for the following ionic  
compounds a sodium phosphate c  
sodium hydroxide e ammonium chloride  
b magnesium d potassium cyanide 5  
Write formulas for compounds formed  
from these pairs of ions b , 6 Name the

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following compounds ...

## **[MOBI] Chapter 9 Chemical Names And Formulas Answers**

Chapter 9 “Chemical Names and Formulas” ... in parenthesis after the name of the metal (Table 9.2, p.255)  
Predicting Ionic Charges Some of the post-transition elements also have more

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than one possible oxidation state. Tin (II) =  $\text{Sn}^{2+}$  Lead (II) =  $\text{Pb}^{2+}$  Tin (IV) =  $\text{Sn}^{4+}$  Lead (IV) =  $\text{Pb}^{4+}$

## **“Chemical Names and Formulas”**

Chapter 9 Chemical Names and Formulas 211. Section Review.

Objectives. • Determine the charges of monatomic ions by using the periodic

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table and write the names of the ions.

- Define a polyatomic ion and write the names and formulas of the most common polyatomic ions.
- Identify the two common endings for the names of most polyatomic ions.

**05 CTR ch09 7/9/04 3:29 PM Page  
211 NAMING IONS 9**

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Chapter 9 Chemical Names and Formulas 221 Name Date 4 Write formulas for the following ionic compounds a sodium phosphate c sodium hydroxide e ammonium chloride b magnesium d potassium cyanide 5 Write formulas for compounds formed from these pairs of ions b , 6 Name the following compounds ...

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