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## **Acids And Bases Chapter Review**

Acids and bases in aqueous solutions will conduct electricity because they contain dissolved ions.

Therefore, acids and

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bases are electrolytes. Strong acids and bases will be strong electrolytes. Weak acids and bases will be weak electrolytes. This affects the amount of conductivity.

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Acids Bases (i) taste sour (i) taste bitter (ii) are corrosive to metals (ii) feel slippery or soapy (iii) change blue litmus red (iii) change red litmus blue (iv) become less acidic on

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mixing (iv) become less basic on mixing with with bases acids While Robert Boyle was successful in characterising acids and bases he could not

## **Notes ACIDS, BASES AND SALTS**

Bases. A chemical compound that increases the concentration of hydrogen ions in aqueous solution.

Arrhenius acid. A



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substance that increases the concentration of hydroxide ions in aqueous solution. Arrhenius base. Acid molecules attract water molecules which take the.

## **Chemistry Chapter 14 ~Acids and Bases Test Review ...**

CHAPTER 14 REVIEW .  
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SHORT ANSWER

Answer the following

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questions in the space provided. 1. a. Write the two equations that show the two-stage ionization of sulfurous acid in water. stage 1:  
$$\text{H}_2\text{SO}_3(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{H}_3\text{O}^+(\text{aq}) + \text{HSO}_3^-(\text{aq})$$
  
b. Which stage of ionization usually produces more ions? Explain your answer.

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SECTION 1 SHORT  
ANSWER Answer the

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following questions in  
the space provided. 1.

Name the following  
compounds as acids:

sulfuric acid a.  $\text{H}_2\text{SO}_4$

sulfurous acid b.  $\text{H}_2\text{SO}_3$

3 hydrosulfuric acid c.

$\text{H}_2\text{S}$  perchloric acid d.

$\text{HClO}_4$  hydrocyanic

acid e. hydrogen

cyanide 2. H

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The equilibrium  
constant for the  
reaction  $A^- +$

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Chapter 15: Acids and  
Bases Acids and Bases  
Arrhenius Definitions:  
acids - compounds that

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produce an increase in  
[H<sup>+</sup>] when dissolved in  
water bases -  
compounds that  
produce an increase in  
[OH<sup>-</sup>] when dissolved  
in water Lewis

Definitions: acids -  
electron pair acceptors  
bases - electron pair  
donors Brønsted-Lowry  
Definitions: acids - H<sup>+</sup>  
donors

## **Chapter 15: Acids and Bases Acids and Bases**

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1 Organic Acids, Bases,  
and Acid Derivatives

Chapter 14  
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review acids and bases

section 3 short answer

answer the following

questions in the space

provided 1 answer the

following questions

according to the br

nsted lowry definitions

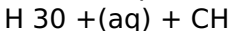
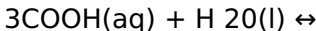
of acids and bases: a

what.

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## **Chapter 14 Review Acids And Bases Answer Key**

ANSWERS Unit 14:  
Review Acids and  
Bases 1) CH



In the equilibrium above,

what are the two

conjugate bases? A. CH

3COOH and H<sub>2</sub>O B. CH

3COO<sup>-</sup> and H<sub>3</sub>CO<sup>+</sup> C. CH

3COOH and H<sub>3</sub>CO<sup>+</sup> D.

CH<sub>3</sub>COO<sup>-</sup> and H<sub>2</sub>O 2)



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Which of the following is the weakest acid in aqueous solution? A.  $\text{C}_6\text{H}_5\text{OH}$   $K_a = 1.3 \times 10^{-10}$  B.  $\text{HCN}$   $K_a = 4.9 \times 10^{-10}$  C.  $\text{H}_2\text{O}$

## **ANSWERS Unit 14: Review Acids and Bases**

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In the Arrhenius definition, acids release protons while in the Bronsted-Lowry definition, they donate protons. Arrhenius said

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that bases donate electron pairs whereas Bronsted and Lowry said that bases accept protons.

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The Lewis definitions of an acid and a base are based on electron pairs, not protons. A Lewis acid is an electron pair acceptor, while a Lewis base is an electron pair donor.

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Use Lewis diagrams to show that  $\text{H}^+(\text{aq}) + \text{OH}^-(\text{aq}) \rightarrow \text{H}_2\text{O}(\ell)$  is an acid-base reaction in the Lewis sense as well as in the Arrhenius and Brønsted-Lowry senses.

## **10.E: Acids and Bases (Exercises) - Chemistry LibreTexts**

This section begins with a review of acids, followed by the following for bases: (1)

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it states the Arrhenius definition of base, (2) it provides you with the information necessary to identify strong and weak bases, and (3) it describes the changes that take place when one weak base (ammonia) dissolves in water (Figure 8.2).

## **Chapter 8 Acids, Bases, and Acid- Base Reactions**

CHAPTER FOURTEEN  
ACIDS AND BASES For



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Review 1. a. Arrhenius acid: produce  $H^+$  in water b. Brønsted-Lowry acid: proton ( $H^+$ ) donor c. Lewis acid: electron pair acceptor The Lewis definition is most general. The Lewis definition can apply to all Arrhenius and Brønsted-Lowry acids;  $H^+$  has an empty  $1s$  orbital and forms bonds to all bases by accepting

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## **CHAPTER FOURTEEN ACIDS AND BASES**

- (14.1) The nature of acids and bases (14.2)
- Acid strength (14.3)
- The pH scale (14.4)
- Calculating the pH of strong acid solutions (14.5)
- Calculating the pH of weak acid solutions (14.6)
- Bases (14.7)
- Polyprotic acids

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Chapter Homework:

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review 1 thru 7.

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review 12, 13. Section

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